

	Molecule	Lewis Structure	Bonded Atoms	Lone Pairs	Shape	P/NP
8.	CO ₂ 4 + (6 × 2) = 16ve ⁻	$\ddot{O} = C = \ddot{O}$	2	0	linear	NP
9.	NO ₃ ⁻ 5 + (6 × 3) + 1 = 24ve ⁻	$\left[\begin{array}{c} \ddot{O} \\ \\ \ddot{O} - N = \ddot{O} \\ \\ \ddot{O} \end{array} \right]^{-1} \leftrightarrow \left[\begin{array}{c} \ddot{O} \\ \\ \ddot{O} = N - \ddot{O} \\ \\ \ddot{O} \end{array} \right]^{-1}$	3	0	trigonal planar	NP
10.	HCN 1 + 4 + 5 = 10ve ⁻	H - C ≡ N:	2	0	linear	P
11.	GaI ₃ 3 + (7 × 3) = 24ve ⁻	$\begin{array}{c} \text{:I:} \\ \\ \text{:I:} - \text{Ga} - \text{:I:} \\ \\ \text{:I:} \end{array} \leftrightarrow \begin{array}{c} \text{:I:} \\ \\ \text{Ga} - \text{:I:} \\ \\ \text{:I:} \end{array}$	3	0	trigonal planar	NP
12.	SO ₃ 6 × 4 = 24 ve ⁻	$\begin{array}{c} \text{:O:} \\ \\ \text{:O:} - \text{S} - \text{:O:} \\ \\ \text{:O:} \end{array} \leftrightarrow \begin{array}{c} \text{:O:} \\ \\ \text{:O:} = \text{S} - \text{:O:} \\ \\ \text{:O:} \end{array}$	3	0	trigonal planar	NP
13.	CO ₃ ²⁻ 4 + (6 × 3) + 2 = 26ve ⁻	$\left[\begin{array}{c} \text{:O:} \\ \\ \text{:O:} - \text{C} = \text{:O:} \\ \\ \text{:O:} \end{array} \right]^{-2} \leftrightarrow \left[\begin{array}{c} \text{:O:} \\ \\ \text{C} - \text{:O:} \\ \\ \text{:O:} \end{array} \right]^{-2}$	3	0	trigonal planar	NP
14.	NCI ₃ 5 + 21 = 26	$\begin{array}{c} \text{:Cl:} \\ \\ \text{:Cl:} - \text{N} - \text{:Cl:} \\ \\ \text{:Cl:} \end{array}$	3	1	trigonal pyramid	P

	Molecule	Lewis Structure	Bonded Atoms	Lone Pairs	Shape	P/NP
22.	PCl_2F_3 $5 + (7 \times 5)$ 40ve^-		5	0	trigonal bipyramid	P
23.	SF_2 $6 + (7 \times 2)$ 20ve^-		2	2	bent	P
24.	XeF_4 $8 + (7 \times 4)$ $8 + 28$ $= 36\text{ve}^-$		4	2	square planar	P